SCH 4C **Molarity of Solutions**

*(Concentration Expression)*

**Molarity** is a quantitative way that chemists usually express the **concentration** of a solution. (A *less ambiguous method*).

Molarity is the concentration of a solution that contains one mole of solute per litre of solution.

**moles of solute**

**Molarity =**

**volume of solution (litres)**

**n (solute)**

**M = ; *mol/L, mmol/mL, kmol/m3***

**V(L)**

Ex: Explain how to prepare 500.0 mL of a 0.50 M solution of oxalic acid (H2C2O4 . 2 H2O).